**Three types of bonds**

1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ : electrons are transferred from one atom to another
2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_: electrons are shared between two atoms
3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_: electrons are shared among many atoms

**Covalent bonds**

* Atoms with incomplete outer shells sometimes \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ their outer electrons. When they do, the \_\_\_\_\_\_\_\_\_\_\_\_\_ number of electrons in their outer shells – their own and those they “\_\_\_\_\_\_\_\_\_\_\_” from other atoms- is often equal to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, a complete shell.

**Ionic Bonds**

* When electrons are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ from one atom to another, they become \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charged. Because of their \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charges, they are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ to each other like a magnet.

